How-To: Fourier Transform Infrared Spectroscopy (FTIR)

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Beer-Lambert Law

$$T = \frac{I}{I_0} = e^{-\varepsilon bc}$$
$$A = -ln(T) = -ln\left(\frac{I}{I_0}\right)$$
$$A = \varepsilon bc$$

T = transmittance (relative amount of light passing through a sample)

I, I_0 = intensity (Wm⁻²) of transmitted and incident radiation, respectively

 ε = molar absorptivity (m²mol⁻¹ or M⁻¹cm⁻¹) or how strongly a chemical species absorbs light at a given wavelength

b = path length (distance light travels through the chemical species)

c = molar concentration of chemical species

A = absorbance (relative amount of light absorbed by sample)

- Absorbance is determined in the IR spectrum, and concentration should be known for a liquid sample
- Spacers of accurately determined thickness can be used for path length, and molar absorptivity can then be calculated
- If molar absorptivity and concentration can be determined, the thickness of a film can be calculated by solving for the path length



Sample Preparation





KBr Salt Plates

- Handled via edges (fragile) with gloves
- No fingerprints or water exposure, clean copiously with ethanol or acetone
- Collect a background spectrum with clean
 KBr plates secured in sample holder
- Place a few drops of sample solution on one salt plate, press the plates together, and collect a sample spectrum
- Use a (clean) spacer between the salt plates with sample when the path length must be known









e







Further Processing

With the CSV files, spectra can be opened in Origin for thorough peak fitting and analysis

